

## Science 9-Chemistry

Worksheet 3-1Chemical Formulas and Names

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Due Date _____
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1. Find the numbers of each kind of atom in the following formulas. The first one is done as an example.

a)  $\text{Ca}(\text{NO}_3)_2$     1 Ca,    2 N's    6 O's

b)  $\text{AlCl}_3$       \_\_\_ Al's    \_\_\_ Cl's

c)  $(\text{NH}_4)_2\text{SO}_4$     \_\_\_ N's    \_\_\_ H's    \_\_\_ S's    \_\_\_ O's

d)  $\text{Al}_2(\text{CO}_3)_3$     \_\_\_ Al's    \_\_\_ C's    \_\_\_ O's

2. Using Combining Capacities, write the correct **formulas** for the following compounds  
The first one is done as an example.

a) calcium fluoride .....Answer    **CaF<sub>2</sub>**

b) sodium iodide .....Answer \_\_\_\_\_

c) potassium sulphide .....Answer \_\_\_\_\_

d) zinc oxide .....Answer \_\_\_\_\_

e) barium chloride .....Answer \_\_\_\_\_

f) rubidium oxide .....Answer \_\_\_\_\_

- g) strontium sulphide.....Answer \_\_\_\_\_
- h) aluminum iodide.....Answer \_\_\_\_\_
- i) iron (III) bromide .....Answer \_\_\_\_\_
- j) copper (II) sulphide .....Answer \_\_\_\_\_
- k) magnesium nitrate .....Answer \_\_\_\_\_
- l) calcium carbonate.....Answer \_\_\_\_\_
- m) silver sulphate.....Answer \_\_\_\_\_
- n) ammonium fluoride.....Answer \_\_\_\_\_
- o) nickel (III) hydroxide .....Answer \_\_\_\_\_
- p) potassium dichromate.....Answer \_\_\_\_\_
- q) sodium cyanide.....Answer \_\_\_\_\_
- r) lead (II) chloride.....Answer \_\_\_\_\_
- s) cobalt (II) sulphate .....Answer \_\_\_\_\_
- t) mercury (I) sulphite.....Answer \_\_\_\_\_
- u) chromium (III) sulphate .....Answer \_\_\_\_\_

3. Write the correct **name** for each of the following compounds: The first one is done as an example.

- a)  $\text{Na}_2\text{S}$  ..... sodium sulphide
- b)  $\text{K}_2\text{SO}_4$  ..... \_\_\_\_\_
- c)  $\text{LiI}$ ..... \_\_\_\_\_
- d)  $\text{BaO}$ ..... \_\_\_\_\_
- e)  $\text{BeBr}_2$ ..... \_\_\_\_\_
- f)  $\text{NaOH}$  ..... \_\_\_\_\_
- g)  $\text{KMnO}_4$  ..... \_\_\_\_\_
- h)  $\text{Be}(\text{NO}_3)_2$  ..... \_\_\_\_\_
- i)  $\text{NH}_4\text{Br}$ ..... \_\_\_\_\_
- j)  $\text{Zn}_3(\text{PO}_4)_2$ ..... \_\_\_\_\_
- k)  $\text{Al}(\text{OH})_3$ ..... \_\_\_\_\_
- l)  $\text{Rb}_2\text{CO}_3$ ..... \_\_\_\_\_
- m)  $\text{Na}_2\text{CrO}_4$ ..... \_\_\_\_\_
- n)  $\text{KClO}_3$ ..... \_\_\_\_\_
- o)  $(\text{NH}_4)_3\text{PO}_4$ ..... \_\_\_\_\_