

Chemistry 11 – Unit 9—Solution Chemistry

Unit Outline

Topic	Activity
Definitions and Types of Solutions	<p>Class Definitions of solute, solvent, solution, saturated, unsaturated and solubility.</p> <p>Class introduction to polar covalent and non-polar covalent solvents and ionic, polar covalent and non-polar covalent solutes.</p> <p>Experiment 16-A—Polar & Non-Polar Solutes and Solvents.</p>
Aqueous Solutions of Ionic and Covalent Compounds	<p>Class Demonstration of Conductivity of Ionic and Covalent Aqueous Solutions.</p> <p>Class Explanation of Conductivity of Ionic and Covalent Aqueous Solutions. See p. 205-206 and p. 209-210 in SW.</p>
Individual Ion Concentrations	<p>Class Explanation of dissociation and concentrations of individual ions in solution.</p> <p>Class Examples of calculations involving dilution, individual ion concentrations and mixtures of solutions.</p> <p>Do Ex. 30-33 and 36 on page 212 of SW.</p>
Precipitation Reactions	<p>Class Demonstration of Precipitation Reactions and Writing Formula, Complete Ionic and Net Ionic Equations for Precipitation Reactions. (See Old Tutorial 27)</p> <p>Experiment 16-D—Precipitation Reactions</p>
Acid-Base Neutralization Reactions and Titrations	<p>Class Examples of Writing Formula, Complete Ionic and Net Ionic Equations for Neutralization Reactions. (See Tutorial 28)</p> <p>Class Explanation of Acid-Base Titrations and Calculations.</p> <p>Hand-In Assignment #15—Precipitation and Neutralization Reactions and Titrations</p>
Review and Test	<p>Short Class Review of Unit 9</p> <p>Test on Unit 9—Solution Chemistry</p>