

Name \_\_\_\_\_

Date \_\_\_\_\_

Due Date \_\_\_\_\_

Mark \_\_\_\_\_/25

*Correct and Hand in Again by \_\_\_\_\_***Chemistry 11****Hand In Assignment # 3 – Ionic Formulas and Names *KEY***

1. Determine the correct formula for each of the following compounds (15 marks)
  - a. rubidium sulphide ..... **Rb<sub>2</sub>S**
  - b. manganese(II) chloride ..... **MnCl<sub>2</sub>**
  - c. sodium oxalate ..... **Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub>**
  - d. ammonium dichromate ..... **(NH<sub>4</sub>)<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>**
  - e. zinc nitrate tetrahydrate ..... **Zn(NO<sub>3</sub>)<sub>2</sub>·4H<sub>2</sub>O**
  - f. tungsten(VI) sulphate ..... **W(SO<sub>4</sub>)<sub>3</sub>**
  - g. copper(I) monohydrogen phosphate ..... **Cu<sub>2</sub>HPO<sub>4</sub>**
  - h. manganese(IV) bisulphate ..... **Mn(HSO<sub>4</sub>)<sub>4</sub>**
  - i. calcium acetate ..... **Ca(CH<sub>3</sub>COO)<sub>2</sub>**
  - j. iron(III) carbonate hexahydrate ..... **Fe<sub>2</sub>(CO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O**
  - k. tin(IV) oxide ..... **SnO<sub>2</sub>**
  - l. ammonium dihydrogen phosphate ..... **NH<sub>4</sub>H<sub>2</sub>PO<sub>4</sub>**
  - m. calcium sulphite octahydrate ..... **CaSO<sub>3</sub>·8H<sub>2</sub>O**
  - n. yttrium(III) chromate ..... **Y<sub>2</sub>(CrO<sub>4</sub>)<sub>3</sub>**
  - o. nickel(III) chlorate trihydrate ..... **Ni(ClO<sub>3</sub>)<sub>3</sub>·3H<sub>2</sub>O**

2. Determine the correct name for each of the following compounds. (10 marks)

- a.  $\text{Rb}_3\text{PO}_4$  ..... rubidium phosphate
- b.  $\text{Fe}(\text{OH})_3$  ..... iron(III) hydroxide
- c.  $\text{Cr}_2(\text{SO}_4)_3 \cdot 5\text{H}_2\text{O}$  ..... chromium(III) sulphate pentahydrate  
or chromium(III) sulfate pentahydrate
- d.  $\text{NH}_4\text{HC}_2\text{O}_4$  ..... ammonium hydrogen oxalate or ammonium binoxalate
- e.  $\text{Pb}(\text{ClO})_4$  ..... lead(IV) hypochlorite
- f.  $\text{Ba}(\text{OH})_2 \cdot 10\text{H}_2\text{O}$  ..... barium hydroxide decahydrate
- g.  $\text{Hg}_2(\text{NO}_3)_2$  ..... mercury(I) nitrate
- h.  $\text{Cu}_2\text{SO}_3$  ..... copper(I) sulphite or copper(I) sulfite
- i.  $\text{Ni}_2(\text{Cr}_2\text{O}_7)_3$  ..... nickel(III) dichromate
- j.  $\text{Ta}(\text{ClO}_4)_2$  ..... tantalum perchlorate or tantalum(II) perchlorate