

Empirical to Molecular Formula

eg). A compound has a molar mass of 60.0 g/mole and its emp. formula is CH_2O .

Find the molecular formula

	Empirical		Molecular
Formula	CH_2O	$\times 2 \rightarrow$	$\text{C}_2\text{H}_4\text{O}_2$
Mass	$12.0 + 2(1.0) + 16.0$ 30.0	$\times 2 \rightarrow$	60.0